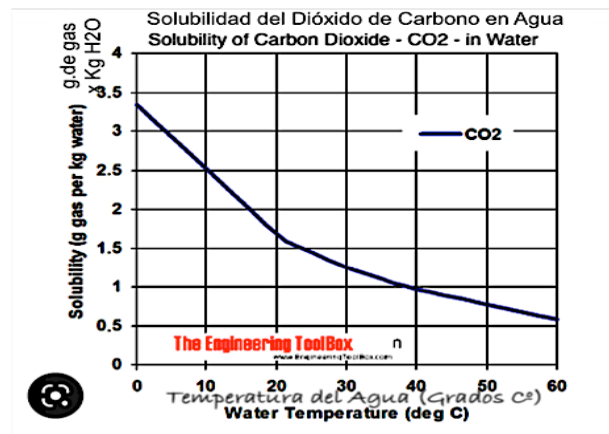




Positive CO₂ capture and its recycling as Urea

Sodium Bicarbonate	Air's Nitrogen	Water	Urea	Na carbonate
	Balance:	214.00	166.00	
2NaHCO ₃ +	2N +	H ₂ O =	CO(NH ₂) ₂ +	Na ₂ CO ₃ +
168.00	28.00	18.00	60.00	106.00
2634.00	439.00	282.21	940.71	1661.93

It is a known fact that if we moisten Baking Soda (Sodium bicarbonate, sodium: a metal) and leave it exposed to air, after a while it will smell like urine (urea). This is because baking soda captures Nitrogen from the air.

Hydrogen the forgotten metal

We usually know H₂CO₃ as Carbonic Acid
just because we think of Hydrogen as a gas and forget that it is a Metal.

Thus:	H ₂ CO ₃	Is Di-hydrogen Carbonate		
The same as	Na ₂ CO ₃	Di-sodium Carbonate		
	Carbonic Gas			
	Local exhausts	Water	Hydrogen carbonate	
	Balance:	62.00	62.00	
	CO ₂ +	H ₂ O =	H ₂ CO ₃	
	44.00	18.00	62.00	
	710.00	290.45	1000.45	

Hydrogen Carbonate	CO ₂ Gas	Water	Hydrogen bicarbonate
	Balance:	124.00	124.00
H ₂ CO ₃ +	CO ₂ +	H ₂ O =	2HHCO ₃
62.00	44.00	18.00	124.00
1000.45	710.00	290.45	2000.91

And, the same as baking soda, NaHCO₃, it captures Nitrogen from the Air

Urea from Hydrogen Bicarbonate

Hydrogen Bicarbonate	Air Nitrogen	Water	Urea	Oxygen
	Balance:	108.00	108.00	
HHCO ₃ +	2N +	H ₂ O =	CO(NH ₂) ₂ +	1.5O ₂
62.00	28.00	18.00	60.00	48.00
2000.91	903.64	580.91	1936.36	1549.09
Urea	Oxygen	Hydrogen Bicarbonate	Nitrogen	Water
Balance:	108.00	108.00		
CO(NH ₂) ₂ +	1.5O ₂ =	HHCO ₃ +	2N +	H ₂ O
60.00	48.00	62.00	28.00	18.00
1936.36	1549.09	2000.91	903.64	580.91